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Application of the SLV-EoS for Representing Phase Equilibria of Binary Lennard-Jones Mixtures Including Solid Phases

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Abstract

In 2003, A. Yokozeki proposed an Equation of State capable of representing solid, liquid, and vapor phases. This equation represents an innovative approach of including solid in phase diagrams. The capability of this Equation of State, named SLV-EoS, of giving a qualitative correct representation of phase diagrams for binary mixtures of Lennard-Jones spheres has been tested in this work. The SLV-EoS has been used for producing the phase diagrams of binary Lennard-Jones mixtures with diameter ratio σ_{11}/σ_{22} ranging from 0.85 to 1, and well-depth ratio $\varepsilon_{11}/\varepsilon_{22}$ ranging from 0.625 to 1.6 at reduced pressure $P^* = P\sigma_{11}^3/\varepsilon_{11} = 0.002$. The obtained phase diagrams have been compared with literature data obtained by Monte-Carlo simulation. The comparison shows the incapability of the SLV-EoS with null binary interaction parameters of predicting solid-liquid azeotrope and eutectic in the investigated range of σ_{11}/σ_{22} and $\varepsilon_{11}/\varepsilon_{22}$. Binary interaction parameters have been regressed to allow the SLV-EoS giving a qualitative representation of the three types of phase diagrams: solid solution, solid azeotrope and simple eutectic. Binary interaction parameters are shown being smoothed functions of σ_{11}/σ_{22} and $\varepsilon_{11}/\varepsilon_{22}$, allowing the prediction of the phase behavior for other Lennard-Jones mixtures or real mixtures of simple fluids.

Keywords: Lennard-Jones fluid, solid-liquid-vapor equilibrium, phase diagram, equation of state, cryogenic processes.

1. Introduction

Solid-liquid phase diagrams are usually classified in six types: solid solution, azeotrope, eutectic with partial immiscibility, eutectic with complete immiscibility, peritectic with eutectic, and molecular compound [1], see Fig. 1.

Matsuoka [2] investigated the frequencies of occurrence of particular types of solid-liquid equilibrium (SLE) diagrams in binary organic mixtures. He found that more than 50% of the systems in the literature exhibit simple eutectic behavior, and about 25% form molecular compounds (type (f) in Fig. 1). About 7% of the studied systems show peritectic with eutectic behavior (type (e) in Fig. 1). The rest of the systems studied exhibit miscibility (partial or total) in the solid phases (type (a), (b), and (c)).

“Classical” modeling approach for calculating solid – liquid equilibrium (SLE), as for instance the one reported by Gmehling and Kolbe [3], assumes solid phases as pure compounds. This assumption is of practical utility in representing types of phase diagrams with total immiscibility in the solid phases. Nevertheless, this approach cannot be used for systems presenting partial or total miscibility in the solid phases. Systems composed of “simple” molecules, like methane, nitrogen, oxygen, noble gases, etc., present generally partial or total miscibility in the solid phases. These systems are of particular interest in cryogenic industrial processes, as for instance air distillation or natural gas purification. Because of the low temperatures at which these processes occur, solidification of impurities is a common phenomenon. Knowledge of phase diagrams for these systems is then important for process design and optimization, allowing avoiding solid deposits. These last can cause energy losses or blockages of separation units. Moreover, the accurate knowledge of phase diagrams including solid phases can also be useful

for studying innovative separation processes. Some examples of binary mixtures presenting solid solubility are given in Ref. 4 for the system argon – krypton, and in Ref. 5 for the system argon – methane.

As stated above, the “classical” modeling approach for SLE cannot represent systems with solid miscibility. In [6] Yokozeki proposed a non-cubic analytical EoS for describing the phase behavior of three thermodynamic states (solid, liquid, and vapor) of matter. The equation is named SLV-EoS. This equation is the first introducing a discontinuity in the isothermal P - v behavior of a fluid. This discontinuity accounts for the fluid-solid transition and hinder the existence of a solid-liquid critical point.

In this work, the authors tested the capability of the SLV-EoS in representing the SLE of systems with partial or total miscibility in the solid phase. These phase behaviors are typical of simple molecules involved in cryogenic processes as air distillation or natural gas purification. For these systems, few phase equilibrium data involving solid phases are available, hindering sometimes the possibility of indentifying clearly the type of solid – fluid diagram of these mixtures. At the same time, Lennard-Jones (LJ) fluids can be considered as very good approximations of the real “simple” molecules objective of this study. Phase diagrams including solid phases for binary mixtures of LJ components have been studied using Monte-Carlo simulation by M. H. Lamm, C. K. Hall, and M. R. Hitchcock [7-10].

The objective of this work is testing the applicability of the SLV-EoS to the study of cryogenic processes. For this purpose, the capability of the SLV-EoS of giving a qualitative correct representation of phase diagrams for binary mixtures of Lennard-Jones fluids, considered as good approximations of the molecules involved in cryogenic processes, has been tested. The SLV-EoS has been used for producing the phase diagrams of binary Lennard-Jones mixtures with diameter ratio σ_{11}/σ_{22} ranging from 0.85 and 1, and well-depth ratio $\epsilon_{11}/\epsilon_{22}$ ranging from 0.625 and 1.6 at

reduced pressure $P^* = P\sigma_{11}^3/\varepsilon_{11} = 0.002$. The obtained phase diagrams have been compared with the corresponding results obtained by Monte-Carlo simulation [8]. This comparison allows verifying if the SLV-EoS predicts in a qualitative correct way the phase behavior of binary LJ mixtures varying the ratios $\varepsilon_{11}/\varepsilon_{22}$ and σ_{11}/σ_{22} . If the SLV-EoS parameters can be obtained as functions of the ratios $\varepsilon_{11}/\varepsilon_{22}$ and σ_{11}/σ_{22} , the EoS can be applied to mixtures of real “simple” fluids, like the fluids involved in the cryogenic air distillation process (mainly nitrogen, oxygen, noble gases, and light hydrocarbons), once the values of ε and σ are known. For these fluids, especially noble gases and methane, Lennard-Jones approximation is representative of the real phase equilibrium behavior. Then, SLV-EoS can be applied to mixtures of these real fluids, allowing a prediction of phase equilibrium where experimental data are not available. At the same time, this approach cannot be applied to complex molecules, like waxes and asphaltenes, which differ too much from the mono-atomic Lennard-Jones approximation.

2. SLV-EoS for LJ fluids

The SLV-EoS proposed by A. Yokozeki [6] is reported in Eq. (1). This equation represents solid, liquid, and vapor phases of matter.

$$P = \frac{RT}{v-b} \frac{v-d}{v-c} - \frac{a}{v^2} \quad (1)$$

In Eq. (1), P is the pressure, R is the gas constant, T is the temperature, v the molar volume, c the liquid covolume, b the solid covolume, a the parameter keeping into account the attractive forces among molecules, d represents the volume for which the repulsive term in Eq. (1) is null.

In order to represent the molecular simulation results for LJ fluids, Eq. (1) has been written in terms of reduced variables:

$$P^* = \frac{T^*}{v^* - b^*} \frac{v^* - d^*}{v^* - c^*} - \frac{a^*}{v^{*2}} \quad (2)$$

Where:

$$\begin{aligned} P^* &= \frac{P\sigma^3}{\varepsilon}; & T^* &= \frac{k_B T}{\varepsilon}; & v^* &= \frac{v}{N_A \sigma^3} \\ a^* &= \frac{a}{N_A^2 \varepsilon \sigma^3}; & b^* &= \frac{b}{N_A \sigma^3}; & c^* &= \frac{c}{N_A \sigma^3}; & d^* &= \frac{d}{N_A \sigma^3} \end{aligned} \quad (3)$$

Eq. (1) can easily be obtained from Eq. (2) substituting the expressions for the reduced variables, Eq. (3), and considering $R = N_A k_B$. N_A is the Avogadro number, and k_B is Boltzmann constant. ε and σ are respectively the well depth and the collision diameter in the Lennard-Jones potential. Parameters a^* and b^* are functions of temperature, as in the original SLV-EoS:

$$a^*(T^*) = a_0 + a_1 T^* \exp(-a_2 T^{*n}) \quad (4)$$

$$b^*(T^*) = b_0 + b_1 \exp(-b_2 T^{*m}) \quad (5)$$

where $a_0, a_1, a_2, n, b_0, b_1, b_2, m$ are adjustable parameters. Parameters of the SLV-EoS in reduced variables, Eqs. (2), (4), and (5), have been regressed for reproducing analytically the temperatures and pressures of the critical and triple point of the pure LJ fluid. Temperature and pressure values at the critical point were taken from [11]: $T_c^* = 1.31$ and $P_c^* = 0.126$. Temperature and pressure

values at the triple point were taken from [12]: $T_t^* = 0.692$ and $P_t^* = 1.21 \times 10^{-3}$. Parameters of Eqs. (2), (4), and (5) have been regressed in order to represent saturation, melting, and/or sublimation reduced pressure for a given reduced temperature. Reduced pressure for vapor-liquid equilibrium data of the pure LJ fluid were produced using the equation presented in Reference 11:

$$\ln P^* = 1.2629T^* - \frac{4.8095}{T^*} - \frac{0.15115}{T^{*4}} \quad (6)$$

Reduced pressure for solid-liquid and solid-gas equilibrium data of the pure LJ fluid were taken from Reference 12.

Parameters of Eqs. (2), (4), and (5) are presented in Tab. 1. These parameters have been obtained minimizing the objective function, f_{ob} , presented in Eq. (7).

$$f_{ob} = \frac{\sum_i^{N_{SLE}} [\ln \varphi^S(\pi^*, P_i^*) - \ln \varphi^L(\pi^*, P_i^*)]^2}{N_{SLE}} + \frac{\sum_i^{N_{VLE}} [\ln \varphi^V(\pi^*, P_i^*) - \ln \varphi^L(\pi^*, P_i^*)]^2}{N_{VLE}} + \frac{\sum_i^{N_{SVE}} [\ln \varphi^S(\pi^*, P_i^*) - \ln \varphi^V(\pi^*, P_i^*)]^2}{N_{SVE}} \quad (7)$$

Where φ^S , φ^L , φ^V are fugacity coefficients of the solid, liquid, and vapor phase, respectively. N_{SLE} , N_{VLE} , and N_{SVE} represent the number of data of solid-liquid, vapor-liquid and solid-vapor equilibrium, respectively.

A $P^* - T^*$ graphical comparison between molecular simulation results and the SLV-EoS model, Eq. (2), is presented in Fig. 2 for the pure LJ fluid.

For representing the phase equilibrium of mixtures, the mixing rules in the form proposed in Eqs. (8) - (11) have been used for the parameters a^* , b^* , c^* , and d^* of Eq. (2). In terms of reduced variables, these mixing rules are written as:

$$a^* = \sum_{i=1}^N \sum_{j=1}^N \sqrt{a_i^* a_j^*} (1 - k_{ij}) x_i x_j \quad \text{where } k_{ij} = k_{ji} \text{ and } k_{ii} = 0 \quad (8)$$

$$b^* = \frac{1}{2} \sum_{i,j=1}^N (b_i^* + b_j^*) (1 - m_{ij}) x_i x_j \quad \text{where } m_{ij} = m_{ji} \text{ and } m_{ii} = 0 \quad (9)$$

$$c^* = \frac{1}{2} \sum_{i,j=1}^N (c_i^* + c_j^*) (1 - n_{ij}) x_i x_j \quad \text{where } n_{ij} = n_{ji} \text{ and } n_{ii} = 0 \quad (10)$$

$$d^* = \frac{1}{2} \sum_{i,j=1}^N (d_i^* + d_j^*) (1 - l_{ij}) x_i x_j \quad \text{where } d_{ij} = d_{ji} \text{ and } d_{ii} = 0 \quad (11)$$

In Eqs. (8) – (11), k_{ij} , m_{ij} , n_{ij} , and l_{ij} are binary interaction parameters, representative of the mixture non-ideality.

3. Phase diagrams of binary LJ mixtures

In References 7 to 10, the effect of the variations of the well-depth ratio, $\varepsilon_{11}/\varepsilon_{22}$, and the diameter ratio, σ_{11}/σ_{22} , on the phase diagram of binary Lennard-Jones mixtures was studied. Phase diagrams were produced using Monte Carlo simulation for diameter ratios ranging from 0.85 to 1, and well-depth ratios ranging from 0.625 to 1.6 for a reduced pressure $P^* = 0.002$.

These ranges of $\varepsilon_{11}/\varepsilon_{22}$ and σ_{11}/σ_{22} allow obtaining phase diagrams as (a), (b), and (c) of figure 1.

In this work, the capability of the SLV-EoS for LJ fluids (LJ SLV-EoS), Eq. (2), of reproducing the phase diagrams obtained with Monte Carlo (MC) simulation has been investigated. Obtained results are represented in figures 3 to 9. For each couple of ratios $\varepsilon_{11}/\varepsilon_{22}$ and σ_{11}/σ_{22} , Fig. (a) represents the diagram obtained via MC simulation [8]. In [8], MC simulation results are presented only graphically, and the corresponding numerical values are no more available from the authors; for these reason, these diagrams have been reproduced in this paper reading data points in the graphs presented in reference 8 with the aid of a specific software. Graphs (b) in Figs. 3-9 present the phase diagrams obtained with the LJ SLV-EoS with null binary interaction parameters in Eqs. (8) - (11). Graphs (c) in Figs. 3-9 present the phase diagrams obtained with the LJ SLV-EoS with binary interaction parameters in Eqs. (8) - (11) regressed for reproducing (at least qualitatively) the phase diagrams obtained with MC simulation. The values used for the binary interaction parameters are presented in Tab. 2.

Binary interaction parameters have been regressed minimizing the objective function, f_{ob}^{mix} , presented in Eq. (12). This objective function is based on the comparison of the equilibrium temperatures and compositions for SLE, VLE, SVE, SSLE, and SLVE.

$$f_{ob}^{mix} = f^{SLE} + f^{VLE} + f^{SVE} + f^{SSLE} + f^{SLVE} \quad (12)$$

Where for SLE, f^{SLE} is calculated as:

$$f^{SLE} = \sum_{i=1}^{N_{SLE}} \left\{ 100 \times \left| \frac{T_{i,EoS}^* - T_{i,ms}^*}{T_{i,ms}^*} \right| + \left| x_{i,EoS}^S - x_{i,ms}^S \right| + \left| x_{i,EoS}^L - x_{i,ms}^L \right| \right\} \quad (13)$$

In Eq. (13), N_{SLE} is the number of molecular simulation data used in the regression procedure; T^* is the solid-liquid equilibrium reduced temperature; x^S and x^L are the compositions of the solid phase and of the liquid phase, respectively; indexes EoS and ms indicate solid-liquid equilibrium properties calculated by the LJ SLV-EoS and molecular simulation, respectively. Analogous expressions have been used for f^{VLE} , f^{SVE} , f^{SSLE} , and f^{SLVE} .

It is reminded that the purpose of this work was not showing the capability of the LJ SLV-EoS of representing precisely phase equilibria produced via molecular simulation. Instead, the aim of the work is finding a way for predicting binary interaction parameters (k_{ij} , m_{ij} , n_{ij} , and l_{ij}) for representing phase equilibrium including solid phases of simple molecules. Lennard Jones fluids are good approximations for simple real fluids, as substances involved in the cryogenic air distillation process. The aim of the paper is then finding correlations of the parameters k_{ij} , m_{ij} , n_{ij} , and l_{ij} as function of σ_{11}/σ_{22} and $\varepsilon_{11}/\varepsilon_{22}$ of Lennard Jones fluids. Once these correlations are obtained, they can be applied for predicting binary interaction parameters for mixtures of real (simple) fluids for which Lennard Jones parameters (and then the ratios σ_{11}/σ_{22} and $\varepsilon_{11}/\varepsilon_{22}$) are known.

In figures 3-9, the following reduced variables are used:

$$P^* = \frac{P\sigma_{11}^3}{\varepsilon_{11}} \quad \text{and} \quad T^* = \frac{k_B T}{\varepsilon_{11}} \quad (14)$$

Generic index ij refers to the interaction between fluid i and j , then index ii refers to the fluid i . A source of discrepancy among MC simulation results and LJ SLV-EoS is the representation of the pure fluid phase diagrams. LJ SLV-EoS reproduces VLE as from Ref. 11; SLE and SVE as from Ref. 12. The accuracy in representing pure fluid phase equilibria is shown in Fig. 2. LJ SLV-EoS

calculations for pure fluids phase diagram are not always in good agreement with MC results of reference [8], as shown in Figs. 3-9. This discrepancy is higher at high values of the reduced temperature. In practice, MC results presented in Ref. [8] extrapolated to pure fluids are not in agreement with the pure fluid data presented in References 11 and 12. These last were used in this work for the regression of the parameters of the pure LJ fluid.

For the reader convenience, Table 3 presents the abbreviations used in Figs. 3-9.

Fig. 3 shows the phase diagram for the binary LJ mixture with $\varepsilon_{11}/\varepsilon_{22} = 0.625$ and $\sigma_{11}/\sigma_{22} = 0.95$. The MC simulation phase diagram (a) presents two SLVE temperatures. The LJ SLV-EoS with null binary interaction parameters reproduces qualitatively the MC simulation results, but the SVE region exists in a wider temperature range (b). As explained above, this is because of the different representation of the pure fluid phase equilibrium between MC simulation [8] and reference 11 and 12 (LJ SLV-EoS reproduces phase equilibrium of reference 11 and 12). The LJ SLV-EoS model with regressed binary interaction parameters allows improving the composition of the liquid and solid phases on the SLVE at low temperature (c).

Fig. 4 shows the phase diagram for the binary LJ mixture with $\varepsilon_{11}/\varepsilon_{22} = 1$ and $\sigma_{11}/\sigma_{22} = 0.95$. The MC simulation phase diagram (a) presents a SL azeotrope. The LJ SLV-EoS with null binary interaction parameters gives a solid solution with a very narrow SLE loop (b). In other words, the liquid phase is never stable for temperatures lower than the lowest melting temperature of the pure fluids. The LJ SLV-EoS model with regressed binary interaction parameters, see Tab. 2, allows representing the SL azeotrope.

The MC simulation phase diagram for the binary LJ mixture with $\varepsilon_{11}/\varepsilon_{22} = 1.6$ and $\sigma_{11}/\sigma_{22} = 0.95$ presents two SLVE lines, and is symmetrical to the diagram of Fig. 3a. The LJ SLV-EoS with null binary interaction parameters gives a good prediction of MC simulation data.

The LJ SLV-EoS model with regressed binary interaction parameters improves the values of the composition of the liquid and solid phases on the SLVE at low temperature. Because these diagrams are qualitatively similar to the diagrams in Fig. 3, graphs have not been presented, while binary interaction parameters are presented in Tab. 2.

Fig. 5 shows the phase diagrams for the binary LJ mixture with $\varepsilon_{11}/\varepsilon_{22} = 0.625$ and $\sigma_{11}/\sigma_{22} = 0.9$. The MC simulation phase diagram (a) presents two SLVE lines, with a SL azeotrope at low temperatures. The LJ SLV-EoS with null binary interaction parameters gives a SVE region with a wider extension in temperature and does not represent the SL azeotrope. The LJ SLV-EoS model with regressed binary interaction parameters improves the temperatures at which SLVE exists and the composition of the liquid and solid phases on the SLVE at low temperature (c). Furthermore, the SL azeotrope at low temperatures is represented.

The MC simulation phase diagram for the binary LJ mixture with $\varepsilon_{11}/\varepsilon_{22} = 1$ and $\sigma_{11}/\sigma_{22} = 0.9$ presents a SL azeotrope at low temperatures. The difference between the azeotrope temperature and the melting temperature of the pure solids is much lower than in Fig. 4a, even if the diagram is qualitatively similar. As in the case of Fig. 4b, the LJ SLV-EoS with null binary interaction parameters gives a solid solution phase diagram. The LJ SLV-EoS model with regressed binary interaction parameters allows representing the SL azeotrope. The azeotrope temperature is also in good agreement with the MC simulation results. Figures are not presented for this case, because it is qualitatively similar to the case shown in Fig. 4.

Fig. 6 shows the phase diagram for the binary LJ mixture with $\varepsilon_{11}/\varepsilon_{22} = 1.6$ and $\sigma_{11}/\sigma_{22} = 0.9$. The MC simulation phase diagram (a) presents one S_1S_2LE line and two SLVE lines. The LJ SLV-EoS with null binary interaction parameters predicts a S_1S_2LE at higher temperatures with respect to the MC simulation data. This behavior gives a narrow range of temperature for the

S₁LE region. Furthermore, composition x_2 in solid 2 of the S₁S₂LE is much lower than the corresponding composition calculated by MC simulation. The LJ SLV-EoS model with regressed binary interaction parameters allows improving the representation of both the temperature and the compositions of the three-phase lines (c).

Fig. 7 shows the phase diagram for the binary LJ mixture with $\varepsilon_{11}/\varepsilon_{22} = 0.625$ and $\sigma_{11}/\sigma_{22} = 0.85$. The MC simulation phase diagram (a) presents one S₁S₂LE line and two SLVE lines. Eutectic behavior is present at low temperatures. The LJ SLV-EoS with null binary interaction parameters does not predict the S₁S₂LE and the eutectic (b). The predicted phase diagram is qualitatively correct at high temperatures, but at low temperatures the results are not in agreement with the MC simulation data. The LJ SLV-EoS model with regressed binary interaction parameters allows representing the S₁S₂LE and the eutectic (c), allowing a qualitative correct representation of the phase diagram produced by MC simulation.

Fig. 8 shows the phase diagram for the binary LJ mixture with $\varepsilon_{11}/\varepsilon_{22} = 1$ and $\sigma_{11}/\sigma_{22} = 0.85$. The MC simulation phase diagram (a) presents eutectic behavior. The LJ SLV-EoS with null binary interaction parameters does not predict the eutectic (b); instead, a solid solution is predicted. The LJ SLV-EoS model with regressed binary interaction parameters allows representing the eutectic (c). The S₁S₂LE temperature is in agreement with the one predicted by MC simulation.

Fig. 9 shows the phase diagram for the binary LJ mixture with $\varepsilon_{11}/\varepsilon_{22} = 1.6$ and $\sigma_{11}/\sigma_{22} = 0.85$. The MC simulation phase diagram presents a S₁S₂LE with eutectic behavior and two SLVE lines (a). The LJ SLV-EoS with null binary interaction parameters does not predict the eutectic (b); furthermore it predicts a S₁S₂LE temperature higher than the melting temperature of component 2. For this reason, the liquid phase of the S₁S₂L equilibrium is richer in component 2 than the

solid phase. Instead, the LJ SLV-EoS model with regressed binary interaction parameters allows representing the eutectic (c). The three-phase equilibrium temperatures and compositions are in good agreement with the MC simulation.

The MC simulation phase diagram for the binary LJ mixture with $\varepsilon_{11}/\varepsilon_{22} = 0.625$ and $\sigma_{11}/\sigma_{22} = 1$ presents two SLVE temperatures. The LJ SLV-EoS with null binary interaction parameters reproduces qualitatively the MC simulation results. With respect to this last, the LJ SLV-EoS model with regressed binary interaction parameters allows improving the representation of the composition of the liquid phases on the SLVE line at low temperature. Figures for this case are very similar to the figures presented for the case $\varepsilon_{11}/\varepsilon_{22} = 0.625$ and $\sigma_{11}/\sigma_{22} = 0.95$ (see Fig. 3) and therefore they have not been presented.

The MC simulation phase diagram for the binary LJ mixture with $\varepsilon_{11}/\varepsilon_{22} = 1.6$ and $\sigma_{11}/\sigma_{22} = 1$ presents two SLVE temperatures. The LJ SLV-EoS with null binary interaction parameters reproduces qualitatively the MC simulation results. The use of binary interaction parameters in the LJ SLV-EoS changes only slightly the obtained phase diagram. The phase diagrams concerning this case are qualitatively similar to the diagrams presented in Fig. 3.

Fig. 10 shows the binary interaction parameters used in Figs. 3-9 plotted as function of the ratio $\varepsilon_{11}/\varepsilon_{22}$ and σ_{11}/σ_{22} . In particular, Figs. 10(a), (c), (e), and (g) show the dependence of k_{ij} , m_{ij} , n_{ij} , and l_{ij} from $\varepsilon_{11}/\varepsilon_{22}$ at constant σ_{11}/σ_{22} ; Figs. 10(b), (d), (f), and (h) show the dependence of the same parameters from σ_{11}/σ_{22} at constant $\varepsilon_{11}/\varepsilon_{22}$.

4. Conclusions

The equation of state for solid, liquid, and vapor phases, recently proposed by A. Yokozeki, has been applied to Lennard-Jones fluids. The LJ SLV-EoS has been compared with molecular simulation results for pure fluids and binary mixtures. The representation of pure fluids phase diagram is satisfactory for VLE, SLE, and SVE.

Phase diagrams for binary LJ mixtures have been produced for σ_{11}/σ_{22} ranging from 0.85 to 1, and $\varepsilon_{11}/\varepsilon_{22}$ ranging from 0.625 to 1.6 for a reduced pressure $P^* = 0.002$. In this range of parameters, three types of solid-liquid equilibria are encountered: solid solutions, solid-liquid azeotrope, and eutectic with partial immiscibility. The LJ SLV-EoS with null binary interaction parameters never predicts a liquid phase stable at temperatures lower than the lowest melting temperature of the pure solids for the whole range of σ_{11}/σ_{22} and $\varepsilon_{11}/\varepsilon_{22}$. This condition is necessary for having eutectic or solid-liquid azeotrope. Another interpretation of this behavior is that the solid-solid miscibility is always over-estimated by the LJ SLV-EoS with null binary interaction parameters.

For representing qualitatively the evolution of the MC simulation phase diagrams, all the four interaction parameters were used. The reason is that the SLV-EoS describes the solid phase as a high-density liquid phase, and mixing rules for the solid and the liquid phases have the same form. For this reason, the miscibility of the solid phases is overestimated. For decreasing the mutual solubility of the solid phases, binary interaction parameters must be used. As a consequence, the LJ SLV-EoS is mainly adapted to the representation of the phase equilibrium of small, simple molecules, showing partial or total solubility in the solid phases.

Binary interaction parameters in the LJ SLV-EoS have been obtained as function of $\epsilon_{11}/\epsilon_{22}$ and σ_{11}/σ_{22} , allowing predicting the phase behavior of other LJ mixtures (not used in the parameters regression procedure) only knowing the LJ parameters of the pure components.

At the same time, phase diagrams for mixtures of real (LJ-like) fluids can be qualitatively predicted once the LJ parameters for the fluids are known. This result is also useful for interpreting the type of phase diagram of mixtures of simple molecules when only few data are available. As future work, the procedure developed in this work for estimating binary interaction parameters of the SLV-EoS for simple molecules will be applied for predicting the phase diagrams of mixtures of real fluids.

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Nomenclature

List of symbols

a	Equation of state parameter [$\text{J}\cdot\text{m}^3/\text{mol}^2$]
a_0	Parameter in Eq. (4)
a_1	Parameter in Eq. (4)
a_2	Parameter in Eq. (4)
b	Solid covolume [m^3/mol]

b_0	Parameter in Eq. (5)
b_1	Parameter in Eq. (5)
b_2	Parameter in Eq. (5)
c	Liquid covolume [m^3/mol]
d	Equation of state parameter [m^3/mol]
f	Fugacity
k_B	Boltzmann constant: $1.380648813 \times 10^{-23}$ [J/K]
k	Binary interaction parameter
l	Binary interaction parameter
m	Binary interaction parameter
n	Binary interaction parameter
N	Number of components in the mixture
N_A	Avogadro number: $6.022141793 \times 10^{23}$ [mol^{-1}]
P	Pressure [Pa]
R	Gas constant: $R = N_A \cdot k_B$ [J/(mol·K)]
T	Temperature [K]
v	Molar volume [m^3/mol]

Greek letters

ε	Well depth in Lennard-Jones potential [J]
σ	Collision diameter in Lennard-Jones potential [m]
φ	Fugacity coefficient

Subscript

c	Related to the critical point
i	Relative to the substance i

<i>j</i>	Relative to the substance <i>j</i>
<i>ij</i>	Relative to the interaction between substance <i>i</i> and the substance <i>j</i>
<i>ji</i>	Relative to the interaction between substance <i>j</i> and the substance <i>i</i>
<i>ii</i>	Relative to the self-interaction for the substance <i>i</i> (considered null in this work)
<i>ms</i>	Obtained by molecular simulation
<i>EoS</i>	Obtained by the equation of state (LJ SLV-EoS)
<i>ob</i>	Objective (function)
<i>SLE</i>	Solid-liquid equilibrium
<i>VLE</i>	Vapor-liquid equilibrium
<i>SVE</i>	Solid-vapor equilibrium
<i>t</i>	Related to the triple point
<i>Superscript</i>	
*	Reduced
<i>L</i>	Liquid phase
<i>m</i>	Parameter in Eq. (5)
<i>mix</i>	Relative to the mixture
<i>n</i>	Parameter in Eq. (4)
<i>S</i>	Solid phase
<i>V</i>	Vapor phase

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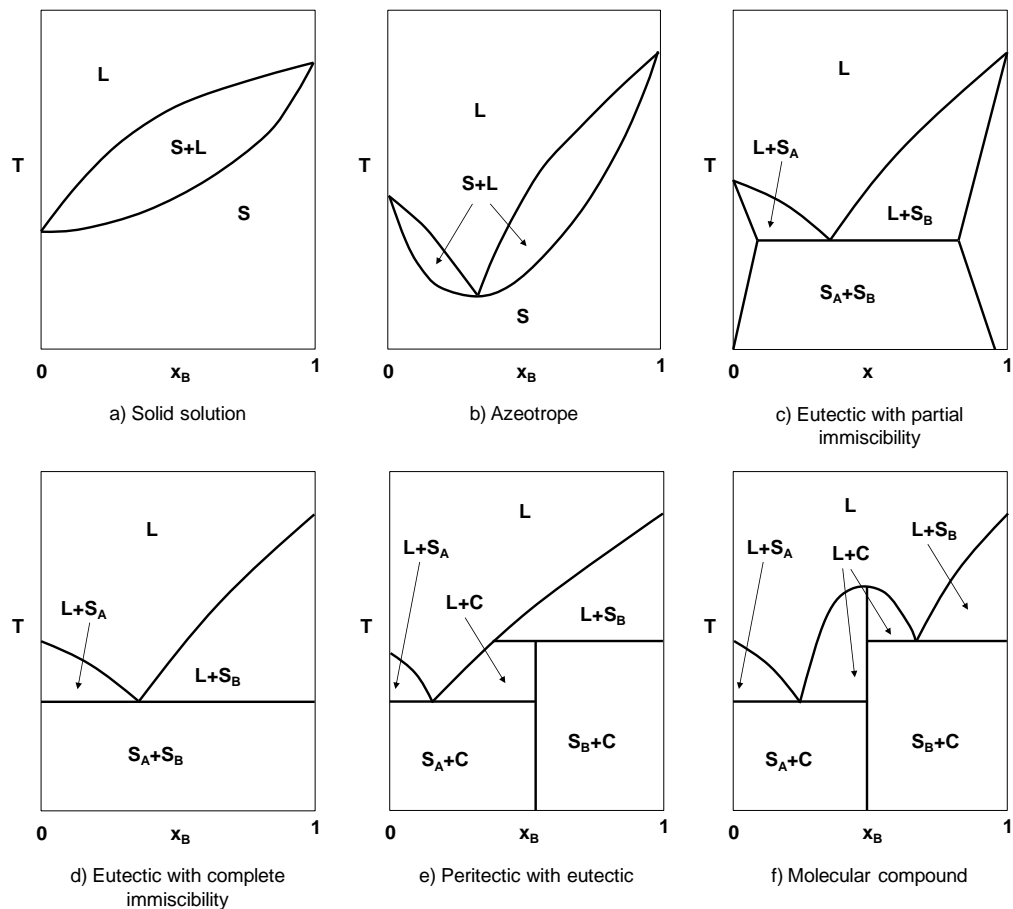


Figure 1. Types of solid–liquid phase diagrams identified by Matsuoka [1]: (a) solid solution, (b) azeotrope, (c) eutectic with partial immiscibility, (d) eutectic with complete immiscibility, (e) peritectic with eutectic, and (f) molecular compound. The phase diagrams are shown in the T - x_B plane where x_B is the mole fraction of component B. L = liquid mixture of A and B, S = solid solution of A and B, S_A = solid solution rich in A, S_B = solid solution rich in B, C = ordered solid with fixed stoichiometric ratio A_mB_n .

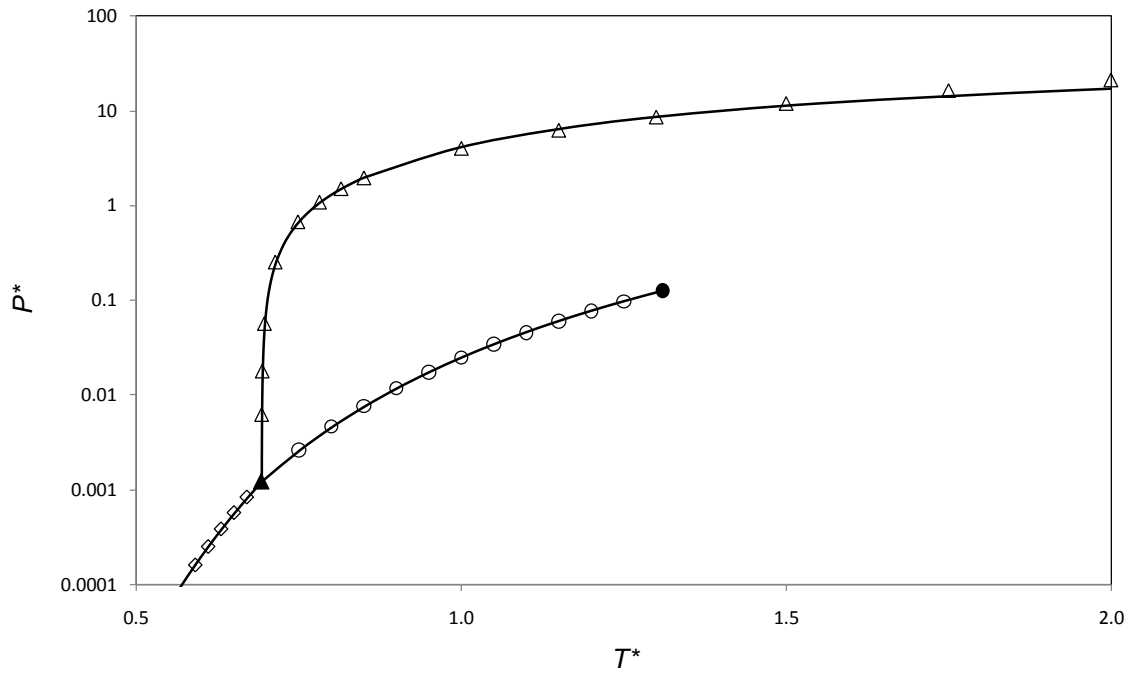


Figure 2. Reduced pressure, P^* , versus reduced temperature, T^* , diagram for the pure Lennard-Jones fluid. Symbols represent molecular simulation results. \circ : VLE [11]; \diamond SVE [12]; \triangle : SLE [12]; \bullet : critical point [11]; \blacktriangle : triple point [12]; —: SLV-EoS model, Eq. (2).

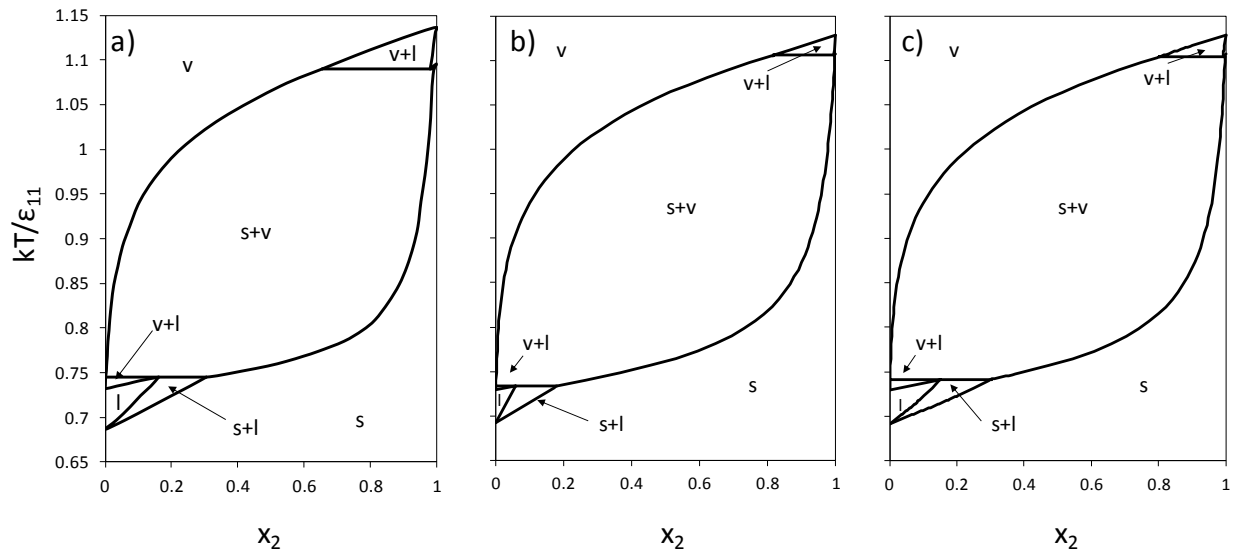


Figure 3. Temperature vs. composition phase diagrams for Lennard-Jones binary mixtures with $\varepsilon_{11}/\varepsilon_{22} = 0.625$ and $\sigma_{11}/\sigma_{22} = 0.95$ at $P^* = 0.002$. (a) MC simulation diagram from [8]; (b) LJ SLV-EoS with null binary interaction parameters; (c) LJ SLV-EoS with regressed binary interaction parameters.

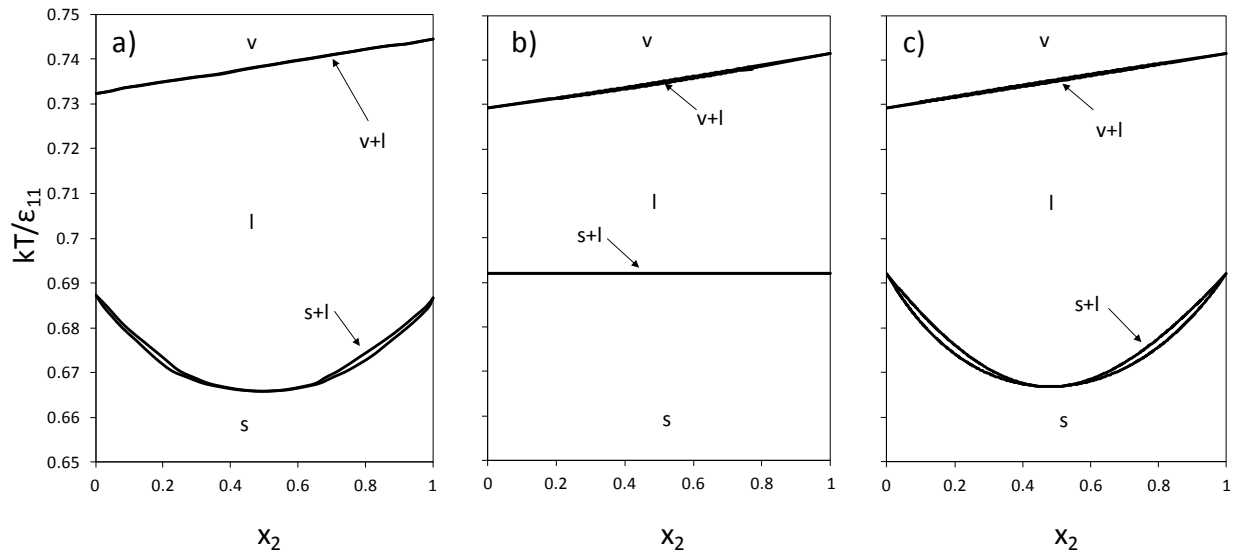


Figure 4. Temperature vs. composition phase diagrams for Lennard-Jones binary mixtures with $\varepsilon_{11}/\varepsilon_{22} = 1$ and $\sigma_{11}/\sigma_{22} = 0.95$ at $P^* = 0.002$. (a) MC simulation diagram from [8]; (b) LJ SLV-EoS with null binary interaction parameters; (c) LJ SLV-EoS with regressed binary interaction parameters.

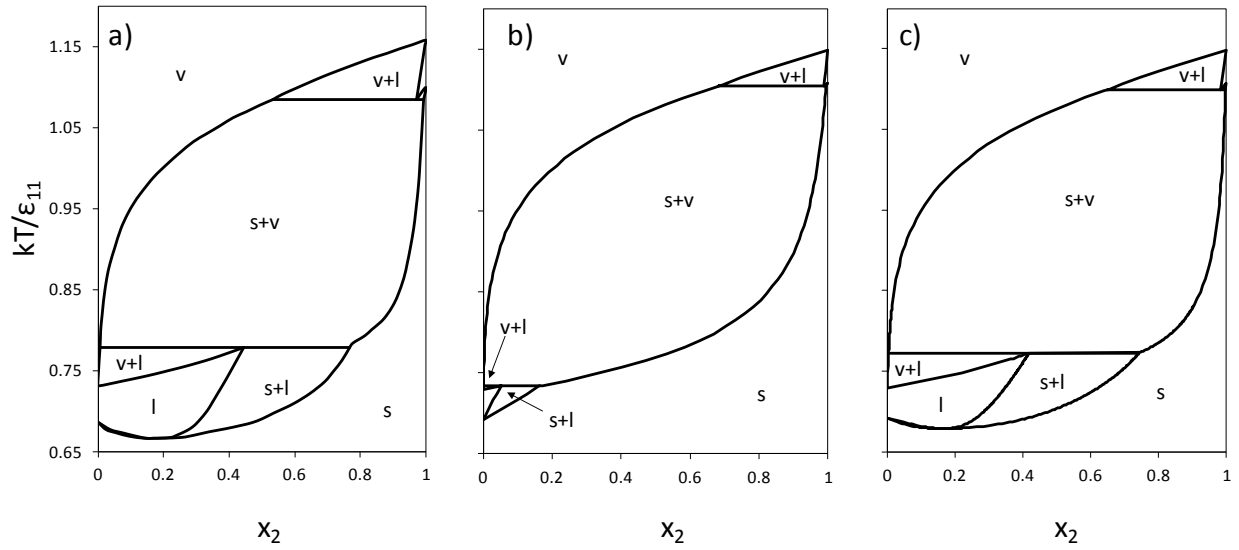


Figure 5. Temperature vs. composition phase diagrams for Lennard-Jones binary mixtures with $\varepsilon_{11}/\varepsilon_{22} = 0.625$ and $\sigma_{11}/\sigma_{22} = 0.9$ at $P^* = 0.002$. (a) MC simulation diagram from [8]; (b) LJ SLV-EoS with null binary interaction parameters; (c) LJ SLV-EoS with regressed binary interaction parameters.

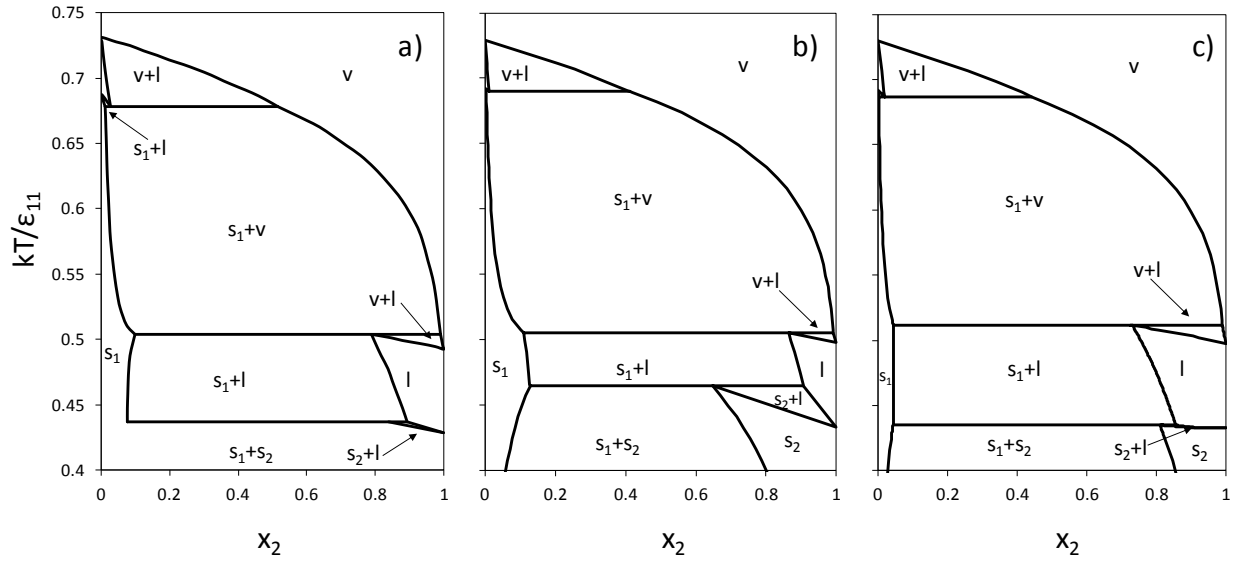


Figure 6. Temperature vs. composition phase diagrams for Lennard-Jones binary mixtures with $\epsilon_{11}/\epsilon_{22} = 1.6$ and $\sigma_{11}/\sigma_{22} = 0.9$ at $P^* = 0.002$. (a) MC simulation diagram from [8]; (b) LJ SLV-EoS with null binary interaction parameters; (c) LJ SLV-EoS with regressed binary interaction parameters.

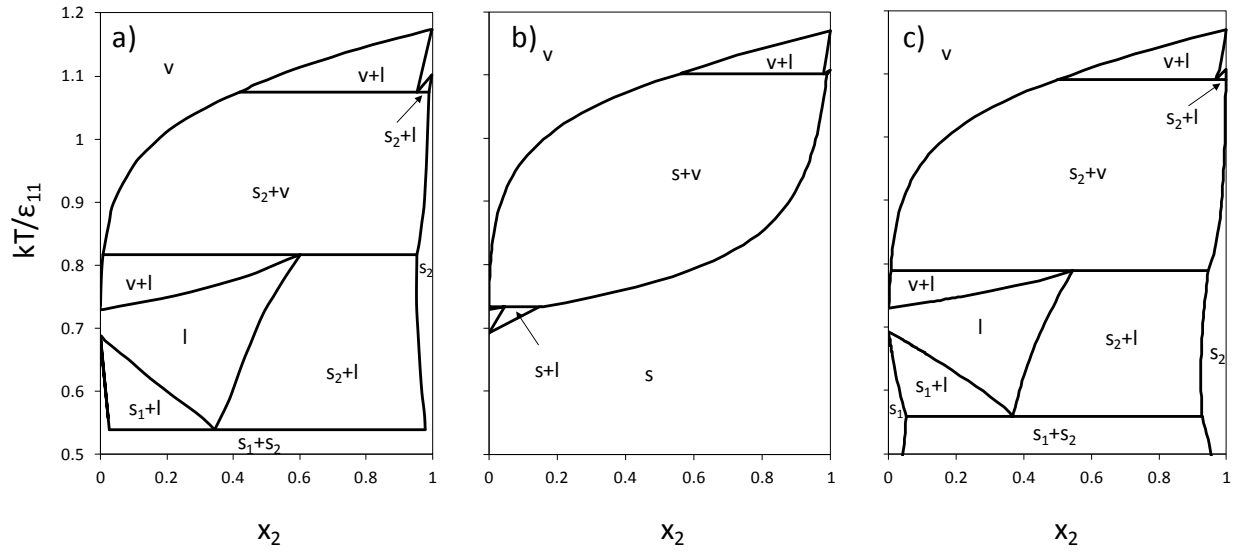


Figure 7. Temperature vs. composition phase diagrams for Lennard-Jones binary mixtures with $\epsilon_{11}/\epsilon_{22} = 0.625$ and $\sigma_{11}/\sigma_{22} = 0.85$ at $P^* = 0.002$. (a) MC simulation diagram from [8]; (b) LJ SLV-EoS with null binary interaction parameters; (c) LJ SLV-EoS with regressed binary interaction parameters.

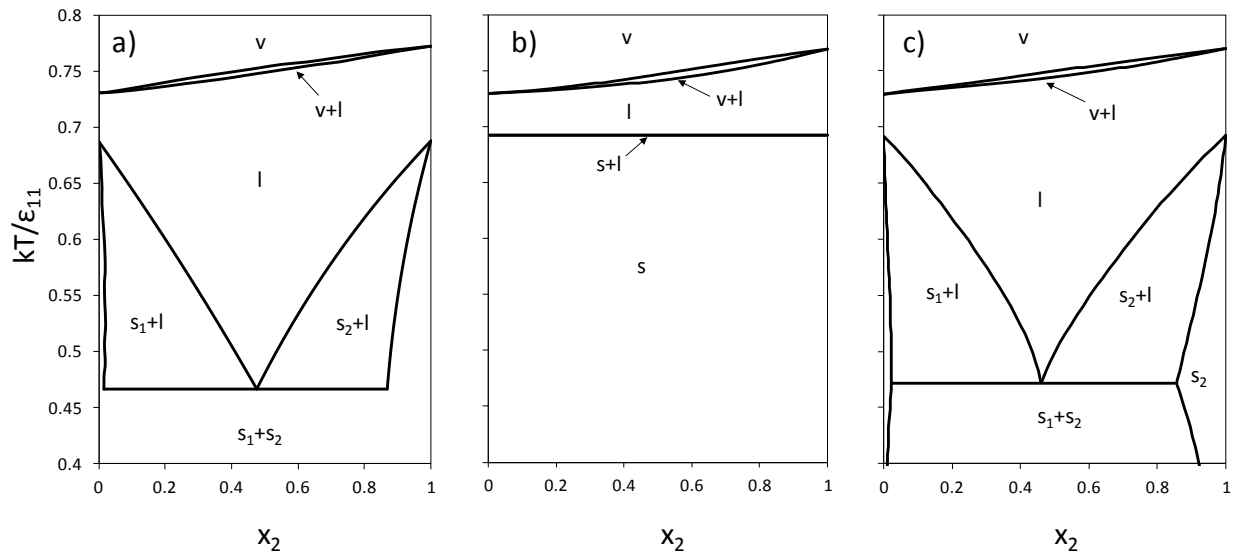


Figure 8. Temperature vs. composition phase diagrams for Lennard-Jones binary mixtures with $\epsilon_{11}/\epsilon_{22} = 1$ and $\sigma_{11}/\sigma_{22} = 0.85$ at $P^* = 0.002$. (a) MC simulation diagram from [8]; (b) LJ SLV-EoS with null binary interaction parameters; (c) LJ SLV-EoS with regressed binary interaction parameters.

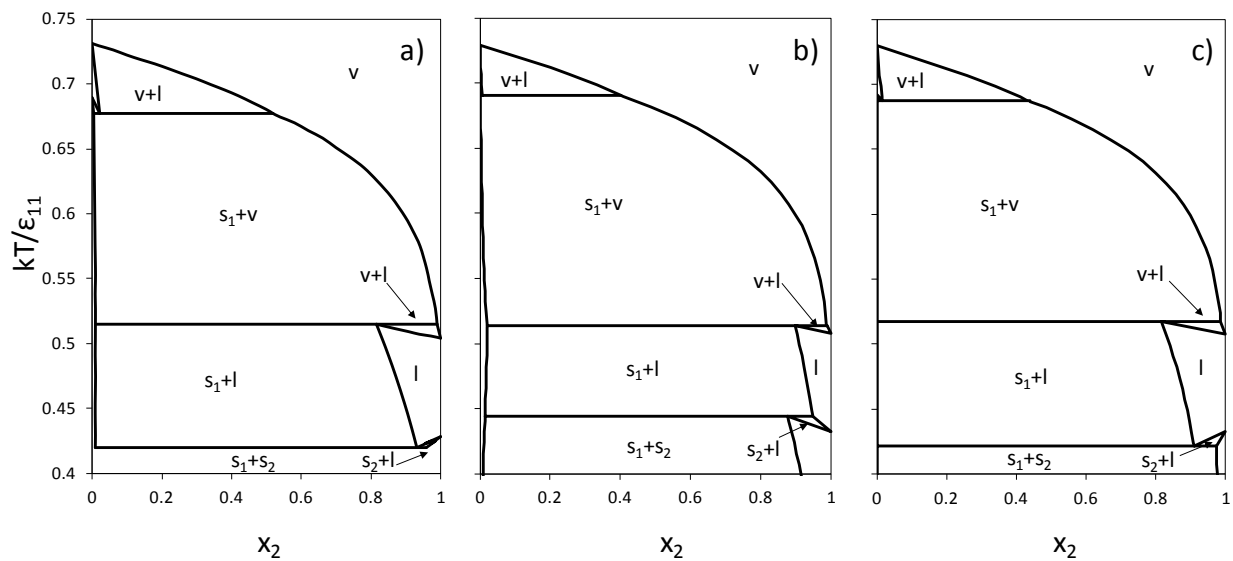


Figure 9. Temperature vs. composition phase diagrams for Lennard-Jones binary mixtures with $\varepsilon_{11}/\varepsilon_{22} = 1.6$ and $\sigma_{11}/\sigma_{22} = 0.85$ at $P^* = 0.002$. (a) MC simulation diagram from [8]; (b) LJ SLV-EoS with null binary interaction parameters; (c) LJ SLV-EoS with regressed binary interaction parameters.

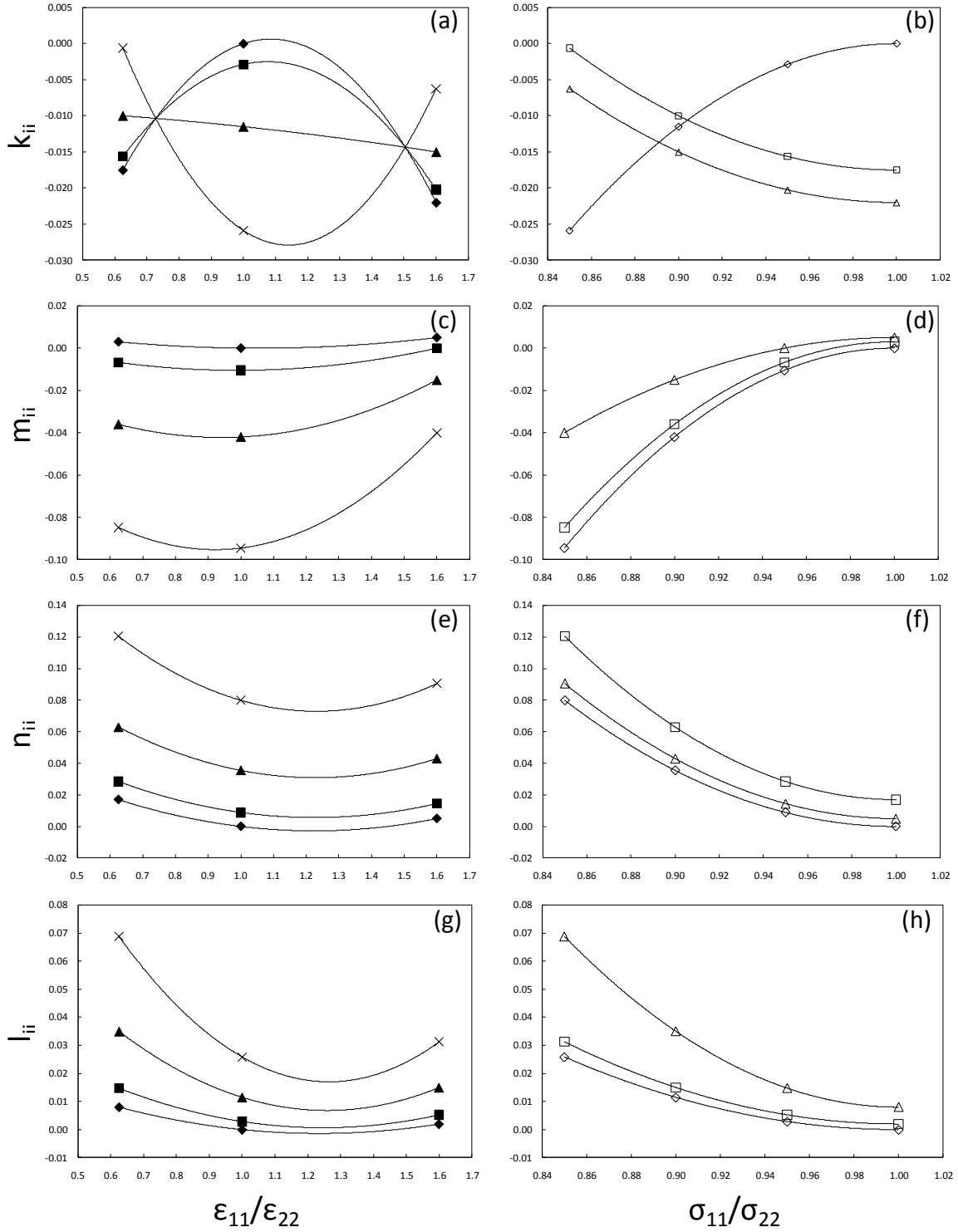


Figure 10. Dependence of the binary interaction parameters k_{ij} , m_{ij} , n_{ij} , l_{ij} from $\epsilon_{11}/\epsilon_{22}$ and σ_{11}/σ_{22} . (—▲—): $\sigma_{11}/\sigma_{22} = 0.9$; (—×—): $\sigma_{11}/\sigma_{22} = 0.85$; (—■—): $\sigma_{11}/\sigma_{22} = 0.95$; (—◆—): $\sigma_{11}/\sigma_{22} = 1$; (—△—): $\epsilon_{11}/\epsilon_{22} = 1.6$; (—◇—): $\epsilon_{11}/\epsilon_{22} = 1$; (—□—): $\epsilon_{11}/\epsilon_{22} = 0.625$; (—): correlations.

Table 1. Parameters of the SLV-EoS in reduced variables, eqs (2), (4), and (5).

Parameter	Value	Units
a_0	2.39647×10^{-1}	-
a_1	4.67098×10^2	-
a_2	4.34036	-
n	3.03527×10^{-1}	-
b_0	1.27853	-
b_1	-3.23646×10^{-1}	-
b_2	1.99173	-
m	1.39554	-
c^*	1.33224	-
d^*	1.29463	-
N_A	$6.022141793 \times 10^{23}$	mol^{-1}
k_B	$1.380648813 \times 10^{-23}$	J/K

Table 2. Binary interaction parameters used for graphs (c) in Figures 3 to 9.

$\varepsilon_{11} / \varepsilon_{22}$	$\sigma_{11} / \sigma_{22}$	$k_{12} = k_{21}$	$l_{12} = l_{21}$	$m_{12} = m_{21}$	$n_{12} = n_{21}$
0.625	0.95	-1.563×10^{-2}	1.475×10^{-2}	-6.75×10^{-3}	2.85×10^{-2}
1	0.95	-2.875×10^{-3}	2.875×10^{-3}	-1.05×10^{-2}	8.875×10^{-3}
1.6	0.95	-2.025×10^{-2}	5.25×10^{-3}	0	1.45×10^{-2}
0.625	0.9	-1×10^{-2}	3.5×10^{-2}	-3.6×10^{-2}	6.3×10^{-2}
1	0.9	-1.15×10^{-2}	1.15×10^{-2}	-4.2×10^{-2}	3.55×10^{-2}
1.6	0.9	-1.5×10^{-2}	1.5×10^{-2}	-1.5×10^{-2}	4.3×10^{-2}
0.625	0.85	-6.25×10^{-4}	6.875×10^{-2}	-8.475×10^{-2}	1.205×10^{-1}
1	0.85	-2.588×10^{-2}	2.5875×10^{-2}	-9.45×10^{-2}	7.9875×10^{-2}
1.6	0.85	-6.25×10^{-3}	3.125×10^{-2}	-4×10^{-2}	9.05×10^{-2}
0.625	1	-1.75×10^{-2}	8×10^{-3}	3×10^{-3}	1.7×10^{-2}
1.6	1	-2.2×10^{-2}	2×10^{-3}	5×10^{-3}	5×10^{-3}

Table 3. Abbreviations used in Figures 3 to 9.

Abbreviations	Subscripts	Meaning
s		solid phase
l		liquid phase
v		vapor phase
s+l		solid-liquid equilibrium
s+v		solid-vapor equilibrium
l+v		vapor-liquid equilibrium
	1	phase rich in component 1
	2	phase rich in component 2